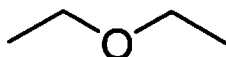
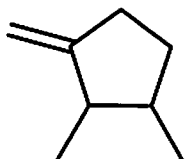
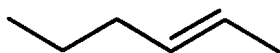
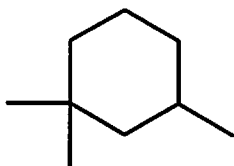
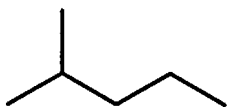
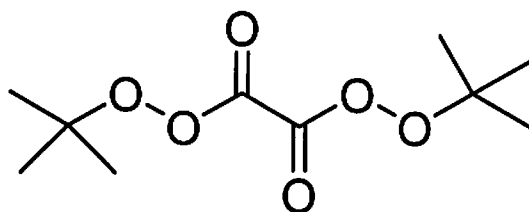


1-5) List the five types of reactions that radicals will typically perform (e.g. *abstraction*).

6-10) Using appropriate arrows, show the mechanisms for the Hydrogen abstraction by  $X^\bullet$  from the *weakest* C-H bond for each of the following molecules (which also appeared in your TWQ1).



11-16) DTBPO is a free radical initiator shown below.

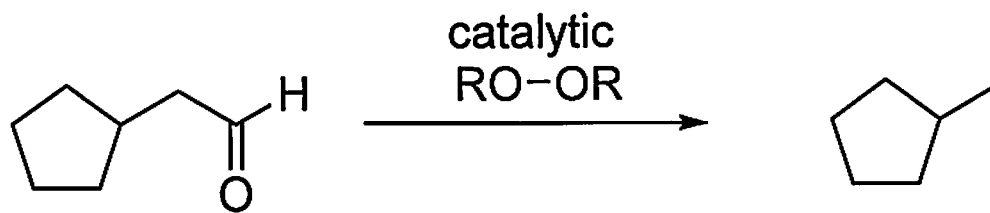


11-13) What do the initials D T B P O stand for? Also explain how each term relates to the above structure.

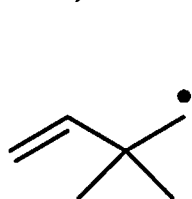
14-16) When heated, DTBPO generates the following species; carbon dioxide, 'butoxy radical, propanone, and methyl radical.

Draw curly arrows to account for the production of these species.

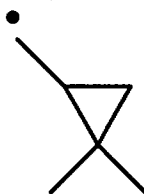
17- 20) Write the mechanism for this free radical chain process, and indicate whether each step is "initiation" or "propagation".



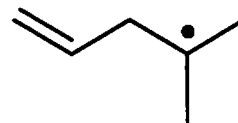
21-25) Consider the isomeric radicals A, B and C shown below:



A



B



C

21) Do you think A or C is more stable?

22) Draw the mechanism for A  $\rightarrow$  B.

23) Is A  $\rightarrow$  B an EXO or ENDO cyclization?

24) If A  $\rightarrow$  B is described as *ring closure*, how would you describe B  $\rightarrow$  C?

25) Draw the mechanism for B  $\rightarrow$  C.

**BONUS POINT:** The above is actually part of a famous kinetic study that illustrated an observation in organic chemistry where large substituents favor ring closure. This effect is named after two chemists – what are their first names?

1-5) List the five types of reactions that radicals will typically perform (e.g. *abstraction*).

Addition

Abstraction

Dimerization

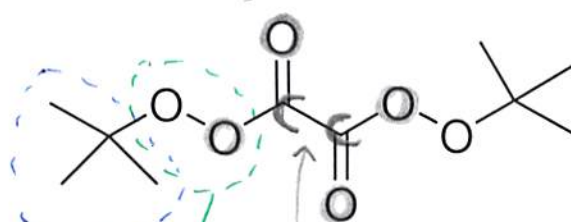
Disproportionation

Fragmentation.

6-10) Using appropriate arrows, show the mechanisms for the Hydrogen abstraction by  $X\cdot$  from the *weakest* C-H bond for each of the following molecules (which also appeared in your TWQ1).



11-16) DTBPO is a free radical initiator shown below.

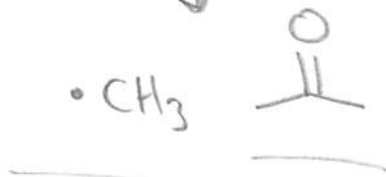
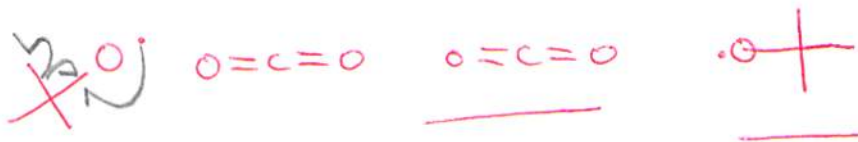
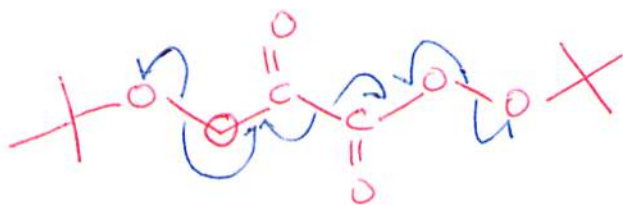


11-13) What do the initials D T B P O stand for? Also explain how each term relates to the above structure.

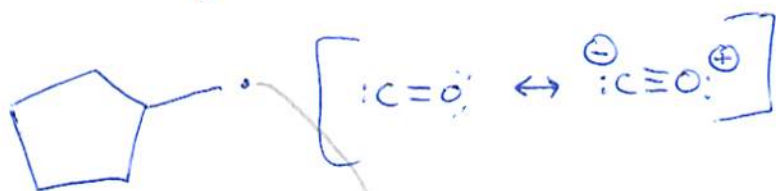
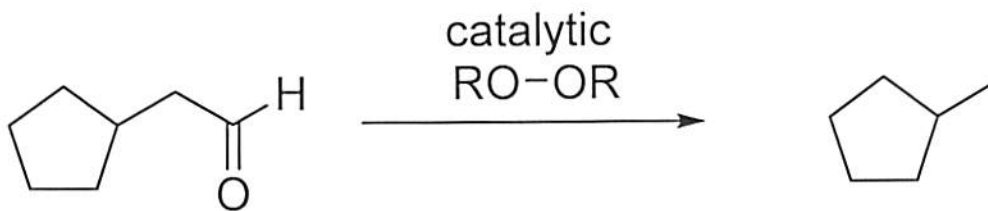
Di Tertiary Butyl Peroxy Oxalate  
Two

14-16) When heated, DTBPO generates the following species; carbon dioxide, 'butoxy radical, propanone, and methyl radical.

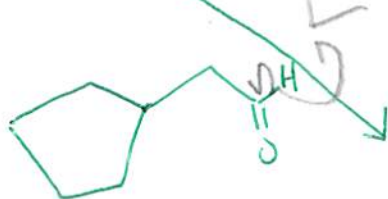
Draw curly arrows to account for the production of these species.



17- 20) Write the mechanism for this free radical chain process, and indicate whether each step is "initiation" or "propagation".



propagation



also propagation

&



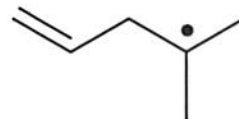
21-25) Consider the isomeric radicals A, B and C shown below:



A



B



C

21) Do you think A or C is more stable?

C

22) Draw the mechanism for A  $\rightarrow$  B.



23) Is A  $\rightarrow$  B an EXO or ENDO cyclization?

24) If A  $\rightarrow$  B is described as *ring closure*, how would you describe B  $\rightarrow$  C?

ring opening

25) Draw the mechanism for B  $\rightarrow$  C.



*BONUS POINT: The above is actually part of a famous kinetic study that illustrated an observation in organic chemistry where large substituents favor ring closure. This effect is named after two chemists – what are their first names?*

Sir Christopher Kelle Ingold & Sir Jocelyn Field Thorpe of

the "Thorpe-Ingold effect"

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